



## An Introduction to Pracical Chemistry; Including Analysis

By John Eddowes Bowman

Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1848 Excerpt: .which gradually becomes purple when exposed to the light. It is nearly insoluble in nitric acid, and considerably less soluble in ammonia than the chloride (429). 434. (C) Perchloride of mercury ( $\text{HgCl}_2$ ) gives a brilliant red precipitate of periclude of mercury ( $\text{HgI}_2$ ), which dissolves in an excess either of the perchloride or of the iodide of potassium. Eg  $\text{Cl}_2 + 2\text{Zn} = \text{HgI}_2 + 2\text{KCl}$ . 435. (C) Starch ( $\text{C}_6\text{H}_{10}\text{O}_5$ ) forms with iodine, even in highly dilute solutions, a dark purple precipitate of iodide of starch. If the iodine is in a state of combination, as in iodide of potassium or hydriodic acid, it is necessary to liberate it before applying the starch; which is readily done by adding a drop or two of nitric acid (iv05) or a solution of chlorine: if nitric acid is employed, a portion of its oxygen combines with the hydrogen or metal with which the iodine was...



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